

1. Give the name of the following polyatomic ions: **(4 pts)**

- a. HCO_3^{1-} b. CN^{1-} c. OH^{1-} d. $\text{C}_2\text{H}_3\text{O}_2^{1-}$

2. Give the name for each element below. **(5 pts)**

- a. Ag
b. Au
c. Pb
d. K
e. Sn

3. Solve using: **heat = S.H x mass X ΔT (4pts)**

How many joules of heat energy are released when 500.0 g of aluminum is heated from 22.0°C to 200.0°C? The specific heat of aluminum is 0.900 J/g°C.

4. Explain the difference between: **(4 pts)**

- a. kinetic energy and potential energy (give examples)

b. metals and non-metals (give examples)

5. List the **SIX remaining** elements that exist as diatomic molecules. Write the name and the formula. The first one is done as an example: **(3 pts)**

Hydrogen H_2